CC13 Groups in the Periodic Table

CC13a Group 1

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| Step | Learning outcome | Had a look | Nearly there | Nailed it! |
| C:\Users\bhuiya_f\Downloads\Steps icons\Steps icons\Progression_icon_L7.jpg | Explain the classification of alkali metals, halogens and noble gases, into groups in the periodic table. |  |  |  |
| C:\Users\bhuiya_f\Downloads\Steps icons\Steps icons\Progression_icon_L6.jpg | Describe the main physical properties of alkali metals. |  |  |  |
| C:\Users\bhuiya_f\Downloads\Steps icons\Steps icons\Progression_icon_L6.jpg | Describe the reactions of lithium, sodium and potassium with water. |  |  |  |
|  | Write word, balanced and H ionic equations (including state symbols) for the reactions of alkali metals. |  |  |  |
|  | Describe the pattern of reactivity of the alkali metals. |  |  |  |
|  | Explain how the electronic configurations of the atoms of alkali metals affect their reactivity. |  |  |  |

CC13b Group 7

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| Step | Learning outcome | Had a look | Nearly there | Nailed it! |
|  | Recall the appearance of chlorine, bromine and iodine at room temperature. |  |  |  |
|  | Describe the trends in colour, melting point and boiling point of chlorine, bromine and iodine down the group, and use these to predict physical properties of other halogens. |  |  |  |
|  | Describe the chemical test for chlorine gas. |  |  |  |
|  | Describe the trends in the reactions of halogens with metals, and use this to predict reactions of other halogens. |  |  |  |
|  | Write word and balanced chemical equations, including state symbols, for the reactions of halogens with metals. |  |  |  |
| C:\Users\bhuiya_f\Downloads\Steps icons\Steps icons\Progression_icon_L6.jpg | Describe hydrogen halides and their chemical properties. |  |  |  |

CC13c Halogen reactivity

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| Step | Learning outcome | Had a look | Nearly there | Nailed it! |
| C:\Users\bhuiya_f\Downloads\Steps icons\Steps icons\Progression_icon_L6.jpg | Describe the relative reactivity of halogens. |  |  |  |
|  | Explain how the reactivity of halogens can be worked out from displacement reactions. |  |  |  |
| D:\WD\Live Job\2016\Sep-16\regcsesciencewordformattingsb3sc13sp4andsp5\Required_Input\Required_Input\TTPP progression steps icons\Progression_icon_L9.jpg | Write balanced chemical equations, including state symbols, for the displacement reactions of halogens. |  |  |  |
|  | HExplain how displacement reactions are examples of redox reactions. |  |  |  |
|  | HWrite ionic equations, including state symbols, for displacement reactions of halogens. |  |  |  |
|  | Explain the order of reactivity of halogens (using electronic configurations). |  |  |  |

CC13d Group 0

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| Step | Learning outcome | Had a look | Nearly there | Nailed it! |
|  | Explain why noble gases are chemically inert by referring to their electronic configuration. |  |  |  |
|  | Describe uses of noble gases linked with their properties. |  |  |  |
| C:\Users\bhuiya_f\Downloads\Steps icons\Steps icons\Progression_icon_L6.jpg | Describe the trends in the physical properties of the noble gases. |  |  |  |
|  | Use trends in physical properties to predict the physical properties of other noble gases. |  |  |  |